

Atomic Structure and Atomic Mass

Name _____

1. What is Z, the atomic number?
2. What is the mass number, A?
3. How many neutrons are in oxygen-17?
4. What is the difference between the mass number and atomic mass?
5. How many protons, neutrons, and electrons in the following isotope? What is the mass number? Identify the element.



6. The isotopic masses and percent abundances are given for element X in the table below. Calculate the weighted average atomic mass. Identify the element.

Isotope	Mass, μ	Percent Abundance
${}^{36}\text{X}$	35.967546	0.3365
${}^{38}\text{X}$	37.962732	0.0632
${}^{40}\text{X}$	39.962383	99.6003