

Periodic Trends

1. Order the elements below from smallest to largest.

I Cl Br F Po

2. Order the ions from smallest to largest

N^{3-} Li^+ O^{2-} F^- Na^+

3. Which of the following are isoelectronic?

Cl P^{3-} Ca^{2+} Rb^+ S^{2-} Br^-

4. Order the following from smallest to largest ionization energy.

Al Cl Mg S Na Si

5. Which of the following would have the highest electron affinity?

Al Cl Mg S Na Si

6. Which would have the highest lattice energy?

Al_2O_3 KCl BaBr_2 LiCl