Periodic Trends

1.	Order	the	elements	below	from	smallest	to	laraest
	0,0,0,	0	0.011101103		,	31110000	~	(3(, 5), 53.5

2. Order the ions from smallest to largest

$$N^{3-}$$
 Li⁺ O^{2-} F⁻ Na^+
Li⁺ < Na^+ < F⁻ < O^{2-} < N^{3-}

3. Which of the following are isoelectronic?

Cl
$$P^{3-}$$
 Ca^{2+} Rb^+ S^{2-} Br^-
Rb⁺ and Br^- are isoelectronic.
 P^{3-} , Ca^{2+} , and S^{2-} are isolectronic.

4. Order the following from smallest to largest ionization energy.

```
Al Cl Mg S Na Si
Na < Mg < Al < Si < S < Cl
```

5. Which of the following would have the highest electron affinity?

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Al Cl Mg S Na Si
```

6. Which would have the highest lattice energy?

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Al_2O_3 KCl BaBr_2 LiCl Al_2O_3
```