Stoichiometry Part 2

Consider the following chemical equation to answer the questions.

$Fe_2O_3\ (s) + CO\ (g) \rightarrow Fe\ (s) + CO_2\ (g)$

a) Balance the equation

b) How many grams of Fe_2O_3 are required to react with 8.75 g of CO? (Don't forget to write the mole ratios)

c) How many grams of solid iron are produced if 8.75 g of CO is reacted?

d) How many grams of solid iron are produced if 10.65 g of Fe_2O_3 reacts with excess CO?