

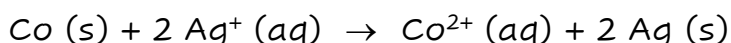
The Activity Series

1. Consider the partial activity series.

Which of the metal(s) can not be oxidized by H^+ ion? **Ag(s)**

Which metal(s) can be oxidized by the H^+ ion? **Co(s), Zn(s), Mg(s), Ba(s)**

Can the following reaction occur? Why?



Yes, because Co (s) can reduce any ion below it on the activity series

2. Which of the following metals will be oxidized by $Pb(NO_3)_2$: Zn, Cu, Fe?

Zn and Fe

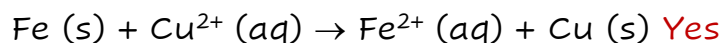
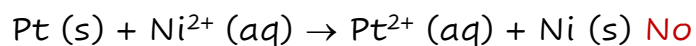
3. Can tin reduce copper (II) ion?

Yes

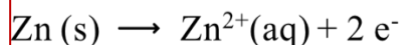
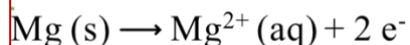
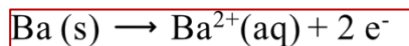
4. Write a reaction for the reduction of lead(II) ion with solid zinc.



5. Which of the following reactions will not occur?



Partial Activity Series



Activity Series of Metals in Aqueous Solution

