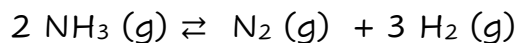


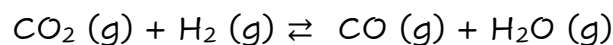
Calculation of K_c/K_p

1. Calculate K_p for the following reaction at 125°C .



The equilibrium pressures are: $P_{\text{NH}_3} = 0.541 \text{ atm}$, $P_{\text{N}_2} = 3.73 \text{ atm}$, and $P_{\text{H}_2} = 11.2 \text{ atm}$

2. Consider the following reaction at 100°C . The initial concentration of $[\text{CO}_2] = 0.325 \text{ M}$ and $[\text{H}_2] = [\text{H}_2] = 0.00768 \text{ M}$. The equilibrium concentration of $[\text{CO}] = 0.0821 \text{ M}$.



a) Write an ICE table.



I				
C				
E				

b) Calculate the equilibrium concentrations of each species.

c) Calculate K_c .

d) Are there more reactants or products at equilibrium?