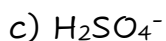


Heats of Formation, ΔH°_f

You can find the standard enthalpies of formation in the appendix of your textbook or at the following link:

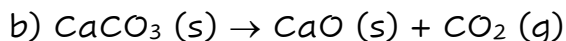
[Standard Enthalpies of Formation](#)

1. Write a balanced formation equation for each the following:



2. Write a balanced formation equation for ethanol, CH_3CH_2OH , and include the value of ΔH°_f . (look in the appendix of your book for ΔH°_f values or use the link at the top of the page).

3. Use the ΔH°_f values in the appendix of your text to determine the enthalpy change for the following reactions: (balance the equations)



4. Use [bond dissociation energies](#) to calculate ΔH°_{rxn} for the following reaction.

