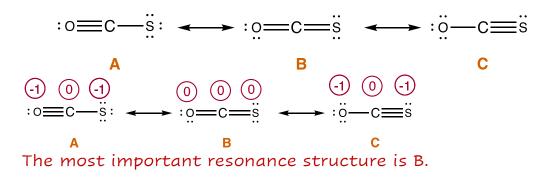
## Lewis Structures Part 3 Formal Charge

1. Draw Lewis structures and show the formal charge for atoms in each of the substances below.



- 2. Draw the best Lewis structure for  $H_3PO_4$ . (Use formal charges)  $H = \bigcup_{i=1}^{n} \bigcup_{j=1}^{n} \bigcup_{i=1}^{n} H_i$  Formal charges are all zero.
  - 3. Consider the following resonance structures. Use formal charges to determine which resonance structure is the most important.



4. Draw the best Lewis structure for ClNO using formal charges.



5. Draw the best Lewis structure for  $H_2SO_3$ .

