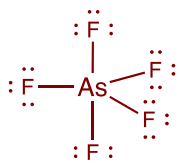


Molecular Polarity

To determine polarity, draw the Lewis structure first. Then determine the electron pair geometry. Remember, the number of charge clouds is equal to the **steric number**.

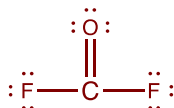
Indicate if the following are polar or nonpolar.

a) AsF_5



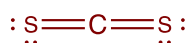
Nonpolar

b) F_2CO



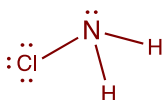
Polar

c) CS_2



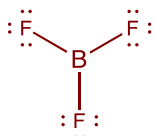
Nonpolar

d) NH_2Cl



Polar

e) BF_3



Nonpolar

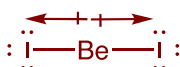
f) N_2O (N is central atom)



Polar

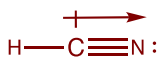
Indicate if the following are polar or nonpolar. Draw the arrows the indicate the bond dipoles.

a) BeI_2



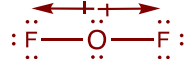
Nonpolar

b) HCN



Polar

c) OF_2



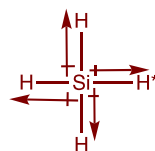
Polar

d) PCl_3



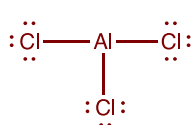
Polar

e) SiH_4



Nonpolar

f) AlCl_3



Nonpolar

Is SF_4 polar or nonpolar? Draw arrows that indicate the bond dipoles.

Polar

