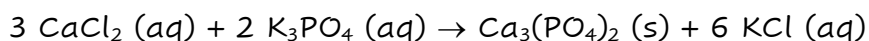


Volumetric (Solution) Stoichiometry and Titration

1. A 45.00 mL sample of HNO_3 was titrated with 0.450 M NaOH. The equivalence point volume was 37.54 mL of NaOH. What is the concentration of the HNO_3 ? (Write a balanced equation)

2. What volume of 0.135 M HClO_4 is required to neutralize 50.00 mL of 0.0926 M $\text{Ba}(\text{OH})_2$? Write a balanced equation.

3. Consider the following chemical equation.



If 25.00 mL of 0.455 M CaCl_2 is mixed with 30.00 mL of 0.365 M K_3PO_4 , how many grams of $\text{Ca}_3(\text{PO}_4)_2$ are formed?